

## CHAPTER 4

### CONCLUSIONS

In this research, unloaded TiO<sub>2</sub> and metal-loaded (metal = Pt, Au and Ag) TiO<sub>2</sub> nanoparticles were successfully synthesized by FSP method. Characterization of these nanoparticles and the sensing films were elucidated and investigated by XRD, BET, SEM with EDS-dot mapping mode and HRTEM. Gas sensing films of all samples were tested towards various flammable gases (H<sub>2</sub>, C<sub>2</sub>H<sub>5</sub>OH and (CH<sub>3</sub>)<sub>2</sub>CO) and environmentally hazardous gases (CO and SO<sub>2</sub>).

#### 4.1 Nanoparticles synthesized by FSP

Unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> (0.25, 0.50, 0.75, 1.0, 2.0 and 3.0 mol%) nanoparticles were successfully synthesized by FSP and structurally characterized by XRD, SEM, EDS and HRTEM. The BET surface area ( $SSA_{BET}$ ) of the nanoparticles was measured by nitrogen adsorption micromeristic technique. From BET measurement, it was found that the calculated particle sizes of all samples were in the same range of ~10–38 nm. XRD patterns revealed that metal-loaded TiO<sub>2</sub> nanoparticles and their corresponding sensing films were crystalline with anatase and rutile phase of TiO<sub>2</sub>. The XRD peaks of Pt and Ag peaks were not found in these patterns, but Au peaks were found in these patterns.

From SEM and EDS data, it can be seen that nanoparticles were spherical in shape, the surface of the nanopowders was porous, and well dispersed without evidence of aggregation. The EDS spectra showed the elemental histograms

corresponding to Ti and O of TiO<sub>2</sub>, Pt, Au and Ag of metal loading. From this observation, it was found that the rough morphology and the rough particle sizes were not changed with increasing Pt, Au and Ag loading levels. HRTEM characterization showed the nanoparticles having clear spherical morphology. The crystallite sizes of spherical unloaded TiO<sub>2</sub> and Pt, Au and Ag-loaded TiO<sub>2</sub> nanoparticles were found to be ranging of 10–80 nm. For Pt, Au and Ag-loaded TiO<sub>2</sub> powder, very small spherical Pt and Ag nanoparticles with diameter of ~2 nm, Au nanoparticles with diameter of ~5–15 nm were found to disperse over the surface of TiO<sub>2</sub> matrix and the presence of Pt, Au and Ag element was confirmed by EDS analysis.

#### 4.2 Comparison of characteristics of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> nanoparticles synthesized by FSP

Table 4.1 shows the summary of characteristics of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> nanoparticles.

**Table 4.1** Summary of characteristics of unloaded TiO<sub>2</sub> and metal-loaded

TiO<sub>2</sub> nanoparticles

Material characterization method	Unloaded TiO <sub>2</sub>	Pt-loaded TiO <sub>2</sub>	Au-loaded TiO <sub>2</sub>	Ag-loaded TiO <sub>2</sub>
XRD	TiO <sub>2</sub> (anatase, rutile)	TiO <sub>2</sub> (anatase, rutile), Pt (cubic face centered)	TiO <sub>2</sub> (anatase, rutile), Au (cubic face centered)	TiO <sub>2</sub> (anatase, rutile), Ag (cubic face centered)

**Table 4.1** (Cont.) Summary of characteristics of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> nanoparticles

Material characterization method	Unloaded TiO <sub>2</sub>	Pt-loaded TiO <sub>2</sub>	Au-loaded TiO <sub>2</sub>	Ag-loaded TiO <sub>2</sub>
BET	Size : 9–10 nm	Size : 14–38 nm	Size : 14–27 nm	Size : 14–24 nm
TEM	Size : ~9–25 nm	Size : ~10–50 nm	Size : ~10–60 nm	Size : ~10–80 nm

#### 4.3 Gas sensing properties

Unloaded TiO<sub>2</sub> film on alumina substrates interdigitated with gold electrodes were prepared by spin-coating technique and tested for gas sensing towards reducing gas such as H<sub>2</sub> and different reducing gases including (CH<sub>3</sub>)<sub>2</sub>CO, CO, C<sub>2</sub>H<sub>5</sub>OH and SO<sub>2</sub>. For the reducing gases, the unloaded TiO<sub>2</sub> synthesized by the FSP method was selective to H<sub>2</sub> with relatively low response ( $S = \sim 2.37$  at 1 vol% H<sub>2</sub>, 400 °C) compared with Pt, Au and Ag-loaded TiO<sub>2</sub> synthesized by FSP. On the other hand, reducing gases like (CH<sub>3</sub>)<sub>2</sub>CO and C<sub>2</sub>H<sub>5</sub>OH, unloaded TiO<sub>2</sub> exhibited (CH<sub>3</sub>)<sub>2</sub>CO and C<sub>2</sub>H<sub>5</sub>OH sensing performances with response ( $S = \sim 19.4$  at 0.20 vol% (350 °C) of (CH<sub>3</sub>)<sub>2</sub>CO and  $S = \sim 21.7$  at 0.10 vol% (400 °C) of C<sub>2</sub>H<sub>5</sub>OH).

Unloaded TiO<sub>2</sub> sensor showed response to reducing gases, H<sub>2</sub>, (CH<sub>3</sub>)<sub>2</sub>CO and C<sub>2</sub>H<sub>5</sub>OH while showed no response to CO and SO<sub>2</sub>. It is well known that the sensor properties can be boosted by activation with some metals such as Pt, Au and Ag. Among various metal tested, Pt, Au and Ag, are the most effective catalysts that can greatly promote sensing of reducing gas including H<sub>2</sub>, CO, (CH<sub>3</sub>)<sub>2</sub>CO and C<sub>2</sub>H<sub>5</sub>OH by

chemical sensitization via “spillover” effect. It can be effectively used to increase response and selectivity as well as to reduce response and recovery times.

For the gas response behavior of Pt, Au and Ag-loaded TiO<sub>2</sub> sensing films several gases have been tested namely H<sub>2</sub>, (CH<sub>3</sub>)<sub>2</sub>CO, CO, C<sub>2</sub>H<sub>5</sub>OH, and SO<sub>2</sub>. The results showed that the gas sensing properties of the metal-loaded TiO<sub>2</sub> sensors were superior to those of the unloaded TiO<sub>2</sub>. Especially, metal-loaded TiO<sub>2</sub> showed higher response, better selectivity, faster response/recovery and better longer term stability usage to especially H<sub>2</sub> than the unloaded TiO<sub>2</sub>. In addition, 2.0 mol% Pt-loaded TiO<sub>2</sub> exhibits the extremely high response of 470 at 1.0 vol% (300 °C) towards H<sub>2</sub>, which much higher than Au and Ag-loaded TiO<sub>2</sub>. In relation to this study the gas sensing properties of materials were related to the surface states and morphology of the material. The gas response could be increased by decreasing the grain size due to high surface/volume ratio.

#### **4.4 Comparison of gas sensing response of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> nanoparticles synthesized by FSP**

Table 4.2 shows the summary of gas sensing performances of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> sensor.

**Table 4.2** Summary of gas sensing performances of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> sensor

Materials	Gas concentration	Response ( $S=R_a/R_g$ )
Unloaded TiO <sub>2</sub> 2.0 mol% Pt-loaded TiO <sub>2</sub> 0.75 mol% Au-loaded TiO <sub>2</sub> 1.0 mol% Ag-loaded TiO <sub>2</sub>	H <sub>2</sub> (1.0 vol%)	~2.37 at 400 °C ~470 at 300 °C ~41.2 at 400 °C ~12.0 at 350 °C
Unloaded TiO <sub>2</sub> 0.50 mol% Pt-loaded TiO <sub>2</sub> 0.75 mol% Au-loaded TiO <sub>2</sub> 0.75 mol% Ag-loaded TiO <sub>2</sub>	(CH <sub>3</sub> ) <sub>2</sub> CO (0.20 vol%)	~19.4 at 350 °C ~244 at 400 °C ~194 at 350 °C ~161 at 350 °C
Unloaded TiO <sub>2</sub> 1.0 mol% Pt-loaded TiO <sub>2</sub> 0.50 mol% Au-loaded TiO <sub>2</sub> Ag-loaded TiO <sub>2</sub>	CO (0.10 vol%)	NO response ~5.39 at 350 °C ~8.60 at 350 °C NO response
Unloaded TiO <sub>2</sub> 0.25 mol% Pt-loaded TiO <sub>2</sub> 0.75 mol% Au-loaded TiO <sub>2</sub> 0.25 mol% Ag-loaded TiO <sub>2</sub>	C <sub>2</sub> H <sub>5</sub> OH (0.10 vol%)	~21.7 at 400 °C ~14.5 at 400 °C ~425 at 350 °C ~20.7 at 400 °C
Unloaded TiO <sub>2</sub> Pt-loaded TiO <sub>2</sub> Au-loaded TiO <sub>2</sub> Ag-loaded TiO <sub>2</sub>	SO <sub>2</sub> (0.05 vol%)	No response No response No response No response

#### 4.5 Suggestions for future work

4.5.1 The gas sensing properties of the unloaded TiO<sub>2</sub> and Pt, Au and Ag-loaded TiO<sub>2</sub> sensors for other gases (such as H<sub>2</sub>S, NH<sub>3</sub>, NO<sub>2</sub>, C<sub>2</sub>H<sub>2</sub> and CH<sub>4</sub>) will be further investigated.

4.5.2 The photocatalytic activity of unloaded TiO<sub>2</sub> and metal-loaded TiO<sub>2</sub> nanoparticles for several organic compounds under UV and/or visible light irradiation will be further investigated as a function of pH.